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PATENT

Amendments to the Specification:

At page 28, please amend lines ^{7 10} ~~11-14~~ as follows:

The preparation of intermediate compounds (XI) and (XII) and other intermediates is described in WO 97/16440-A1, published May 9, 1997 by Janssen Pharmaceutica N.V., which is incorporated ~~disclosed~~ herein by reference as well as in ~~other publications mentioned in WO 97/16440-A1, such as, e.g. EP-0,532,456-A and U.S. Pat. No. 5,310,743.~~
The relevant sections of WO 97/16440-A1 are reproduced below:

After the paragraph at page 28, lines ^{7 10} ~~11-14~~, please insert the following:

The compounds of formula (I) can be prepared by reductively N-alkylating an intermediate of formula (II) with an intermediate of formula (III). Said reductive N-alkylation may be performed in a reaction-inert solvent such as, for example, dichloromethane, ethanol, toluene or a mixture thereof, and in the presence of an appropriate reducing agent such as, for example, a borohydride, e.g. sodium borohydride, sodium cyanoborohydride or triacetoxy borohydride. In case a borohydride is used as a reducing agent, it may be convenient to use a complex-forming agent such as, for example, titanium(IV)isopropylate as described in J. Org. Chem, 1990, 55, 2552-2554. Using said complex-forming agent may also result in an improved *cis/trans* ratio in favour of the *trans* isomer. It may also be convenient to use hydrogen as a reducing agent in combination with a suitable catalyst such as, for example, palladium-on-charcoal or platinum-on-charcoal. In case hydrogen is used as reducing agent, it may be advantageous to add a dehydrating agent to the reaction mixture such as, for example, aluminium *tert*-butoxide. In order to prevent the undesired further hydrogenation of certain functional groups in the reactants and the reaction products, it may also be advantageous to add an appropriate catalyst-poison to the reaction mixture, e.g., thiophene or quinoline-sulphur. Stirring and optionally elevated temperatures and/or pressure may enhance the rate of the reaction.

